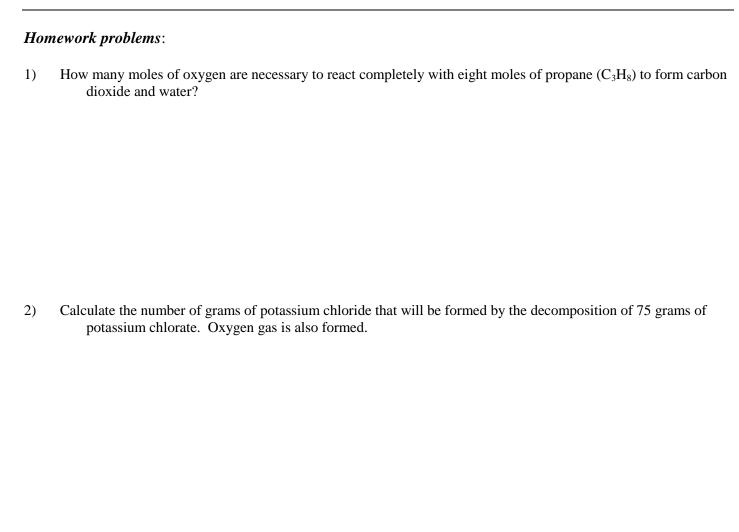
AP CHEMISTRY

TOPIC 2: STOICHIOMETRY, PART A

Day 13:

• Stoichiometry (Basics)



3) How many moles of lead(II) nitrate will be needed to react with sodium chromate to produce 33.2 kg of lead(II)chromate? Sodium nitrate is also produced.

4)	How many grams of hydrogen gas are produced when 45.0 g zinc metal completely replaces hydrogen in hydrochloric acid in a single replacement reaction? Zinc chloride is the other product.
5)	Calculate the mass (in kilograms) of solid gold produced in the single replacement reaction when gold(III) nitrate reacts with 75.0 grams of solid copper to form copper(II) nitrate and solid gold.
6)	Calculate the mass (in milligrams) of decane gas, $C_{10}H_{22}$, when it reacts with 45.0 kilograms of oxygen gas in a combustion reaction.