

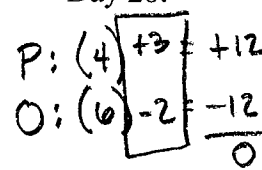
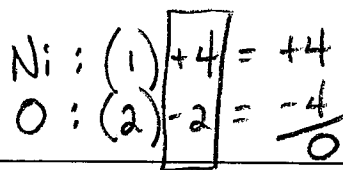
# AP CHEMISTRY

## TOPIC 2: STOICHIOMETRY, PART G

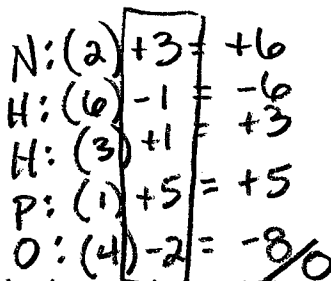
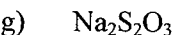
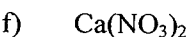
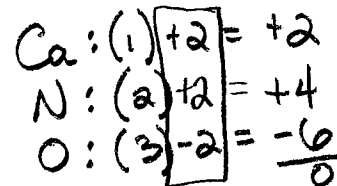
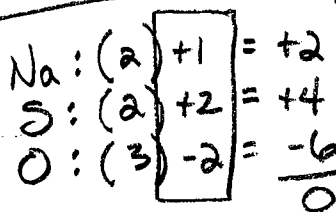
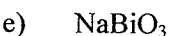
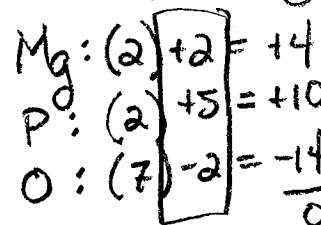
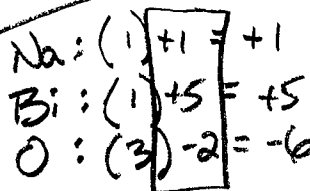
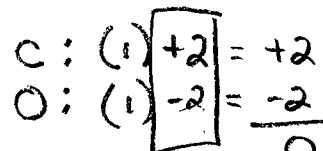
Day 28:

Oxidation / Reduction Equations:

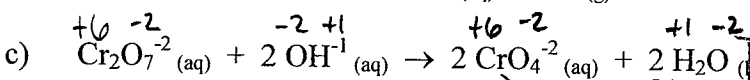
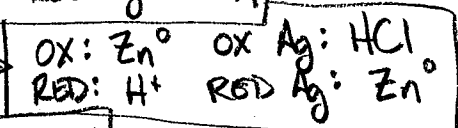
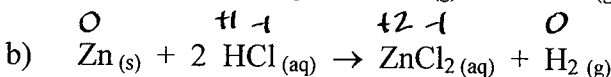
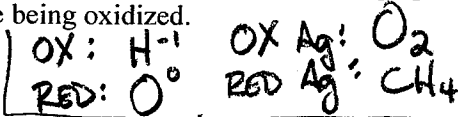
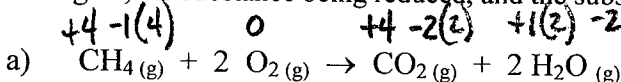
- Oxidation Numbers
- Acid Redox



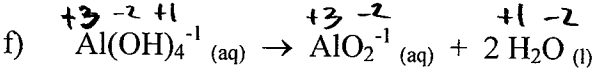
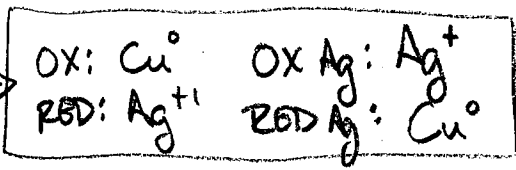
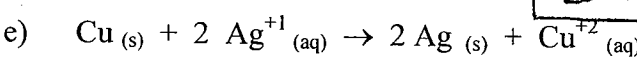
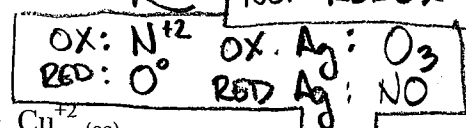
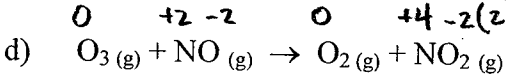
1) Determine the **oxidation number** for each ELEMENT in the chemical formula.



2) Specify which of the following are oxidation-reduction reactions and identify the oxidizing agent, the reducing agent, the substance being reduced, and the substance being oxidized.

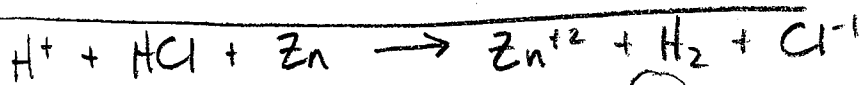
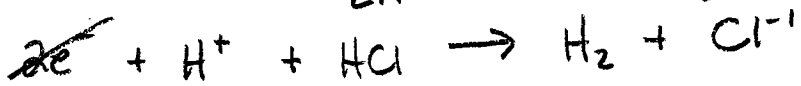
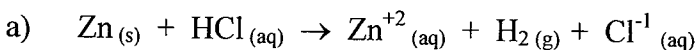


NOT REDOX!



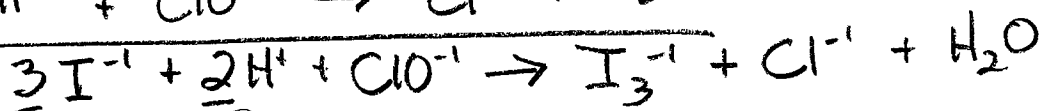
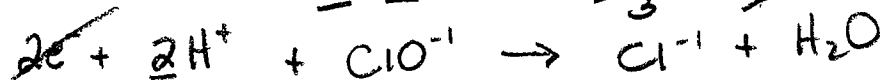
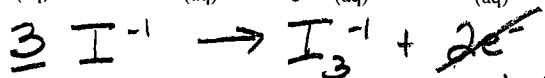
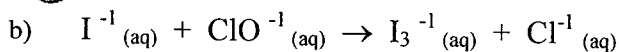
NOT REDOX

3) Balance the following REDOX equations that occur in an acidic solution:



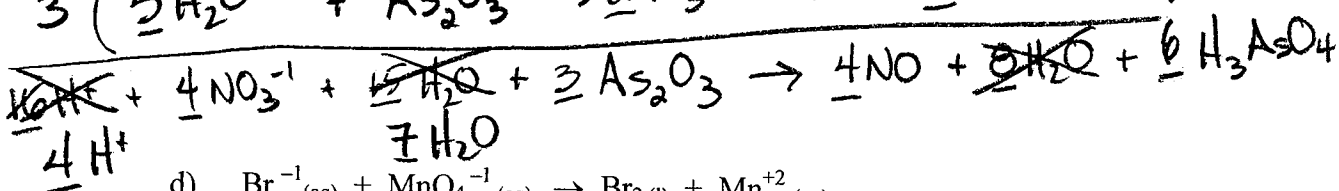
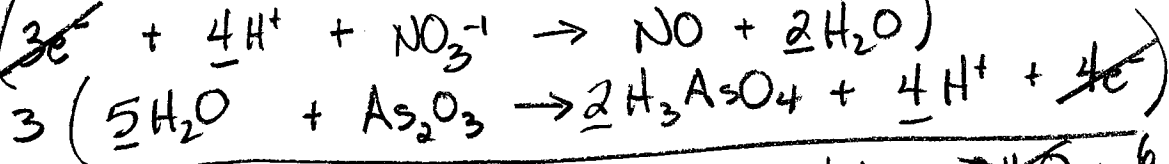
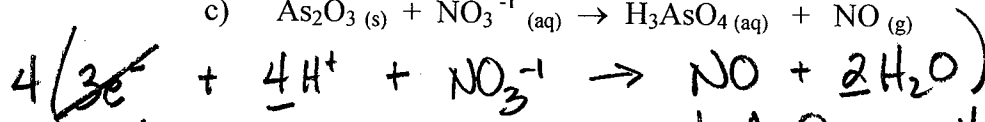
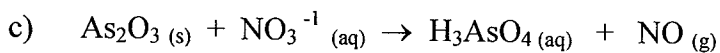
(+1)

(+1)

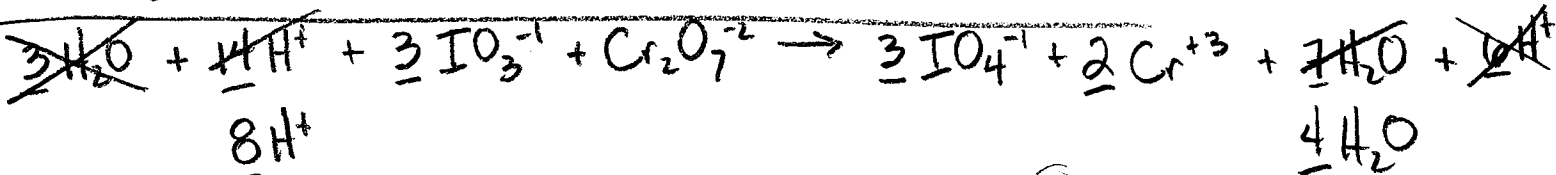
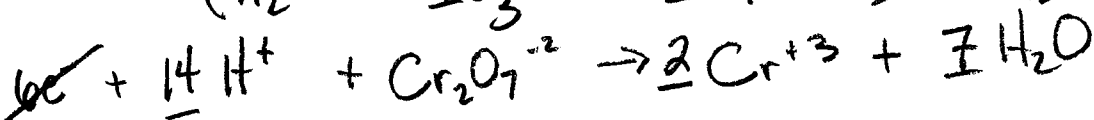
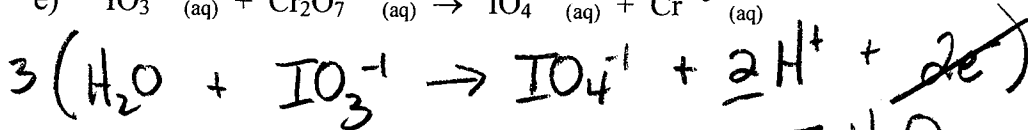
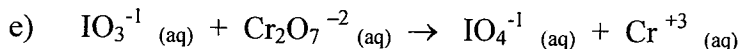
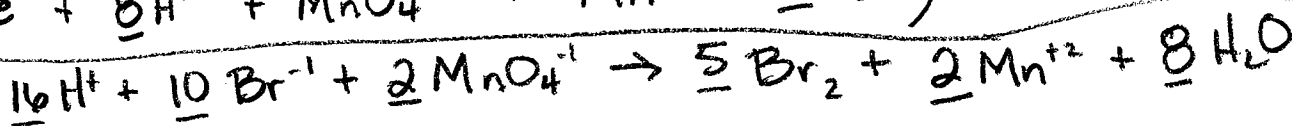
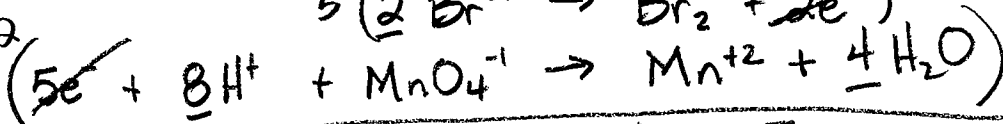
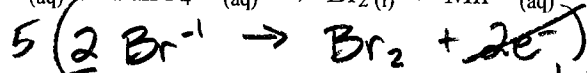
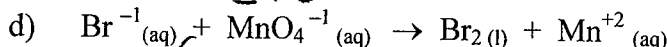


(-2)

(-2)



4H<sup>+</sup>



8H<sup>+</sup>

4H<sub>2</sub>O

(+3)

(+3)