

THE POTENTIAL CHARGE AN ELEMENT MAY HAVE WHEN FORMING AN IONIC BOND WITH ANOTHER ELEMENT

AP CHEMISTRY

TOPIC 2: STOICHIOMETRY, EXAMPLE PROBLEMS

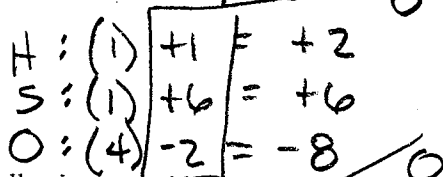
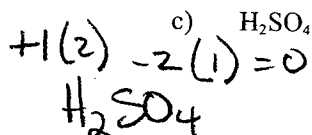
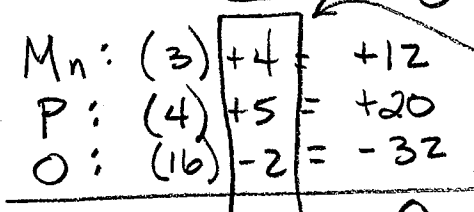
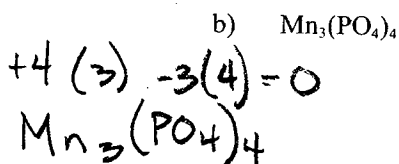
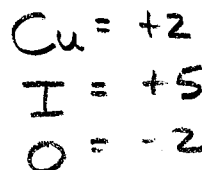
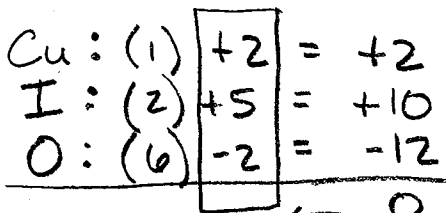
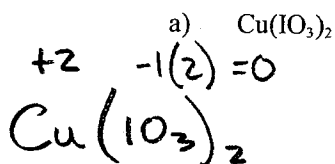
Day 28:

Oxidation / Reduction Equations:

- Oxidation Numbers
- Acid Redox

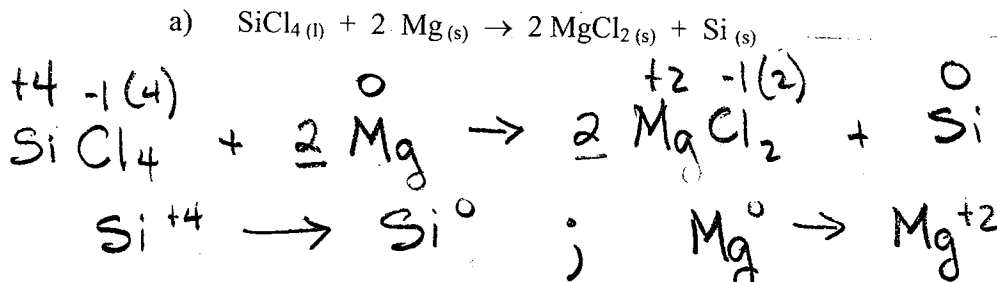
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1) Determine the *oxidation number* for each ELEMENT in the chemical formula

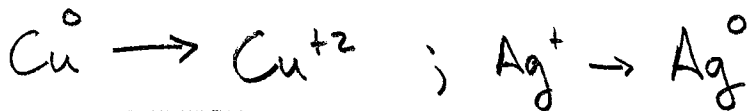
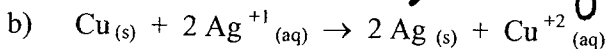


COMPOUND OR ELEMENT THAT "CAUSES ANOTHER" TO OXIDIZE

2) Specify which of the following are oxidation-reduction reactions and identify the oxidizing agent, the reducing agent, the substance being reduced, and the substance being oxidized.

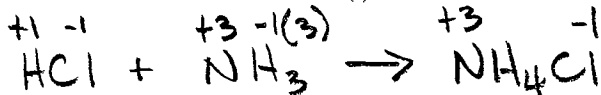
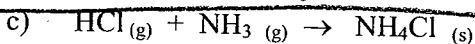


OX: Mg^0
RED: Si^{+4}
OX Ag: SiCl_4
RED Ag: Mg^0



ONLY REACTANTS CAN ANSWER THESE QUESTIONS

OX: Cu^0 ; RED: Ag^+ ; OX Ag: Ag^+ ; RED Ag: Cu^0



SINCE NONE OF THE ELEMENTS CHANGED - THIS REACTION IS NOT REDOX

$3\text{H}^+ \rightleftharpoons 1\text{H}^-$

