AP CHEMISTRY

TOPIC 1: CHEMICAL FOUNDATIONS, PART B

Day 3:

- Early History of Chemistry
- Fundamental Chemical Laws
- Dalton's Atomic Theory
- The atom, and its components

Homework problems:

1) Which of the following is true about an individual atom? Explain your choice.
   a) An individual atom should be considered to be a solid
   b) An individual atom should be considered to be a liquid
   c) An individual atom should be considered to be a gas
   d) An individual atom cannot be considered to be a solid, liquid, or gas.

2) What evidence led to the conclusion that cathode rays had a negative charge (recall from general chem.)

3)

Symbol	Number of Protons in Nucleus	Number of Neutrons in Nucleus	Number of Electrons	Net Charge
	33	42		3+
$^{128}_{52}Te^{2-}$				
	16	16	16	
	81	123		1+
$^{238}_{92}U$				
$^{195}_{78}Pt$				
	20	20		2+
$^{89}_{39}Y$				
	35	44	36	
	15	16		3-
		76	54	1-
$^{120}_{50}Sn^{4+}$				
	6	6		+4