



- Molecules and Ions
 - Periodic Table
 - Naming “simple” compounds
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Homework problems:

- 1) Name or write the formula for the following compounds: “binary molecular compounds”
 - a) S_4N_2
 - b) BI_3
 - c) P_2O_5
 - d) antimony tribromide
 - e) diphosphorus tetroxide
 - f) tetranitrogen dioxide
- 2) Name or write the formula for the following compounds: ionic compounds
 - a) barium fluoride
 - b) calcium nitride
 - c) aluminum oxide
 - d) Li_2O
 - e) AgI
 - f) NaF
- 3) Name or write the formula for the following compounds: ionic compounds (with polyatomics)
 - a) $MgSO_4$
 - b) Na_2SO_2
 - c) NH_4ClO_3
 - d) $Mg(C_2H_3O_2)_2$
 - e) NH_4IO_3
 - f) $AgNO_3$
 - g) magnesium hydroxide
 - h) potassium phosphate
 - i) calcium nitrate
 - j) barium hydroxide
 - k) sodium sulfite
 - l) potassium perchromate
 - m) lithium sulfate

4) Name or write the formula for the following compounds: ionic compounds II

- a) FeS
- b) CoCO_3
- c) CuSO_4
- d) SnBr_4
- e) $\text{Hg}(\text{IO}_3)_2$
- f) PbS
- g) MnO_2
- h) $\text{Au}(\text{NO}_3)_3$
- i) $\text{Sn}(\text{OH})_4$
- j) tungsten(III) oxide
- k) nickel(III) sulfite
- l) mercury(I) sulfide
- m) copper(II) phosphate
- n) bismuth(III) iodate
- o) lead(II) chromate
- p) copper(I) iodide

5) Name or write the formula for the following compounds: acids

- a) nitric acid
- b) hydrofluoric acid
- c) hydrobromic acid
- d) chloric acid
- e) acetic acid
- f) carbonic acid
- g) hydrochloric acid
- h) HCN
- i) H_3P
- j) H_2SO_4
- k) HCl
- l) H_2S
- m) H_3PO_4
- n) HI