

AP CHEMISTRY

TOPIC 5: BONDING, PART A

Day 49:

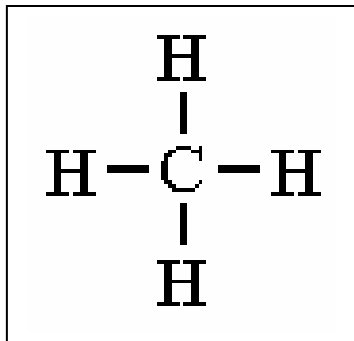
- Covalent Bonds
- Dipole Moment
- Lewis Structures
- Ionic Bonds
- Polarity of Molecule

1) Determine if the molecule has non-polar covalent bonds, polar covalent bonds, or ionic bonds. Then draw the Lewis structure. Then determine if the molecule is Polar (if applicable):

a) CH_4

P.C. Bonds

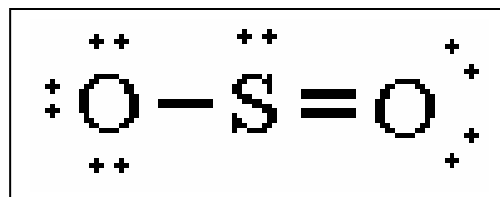
Non-Polar Compound



b) SO_2

P.C. Bonds

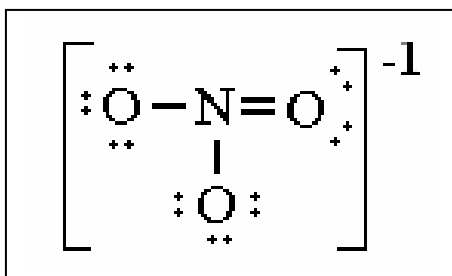
Polar Compound



c) NO_3^{-1}

P.C. Bonds

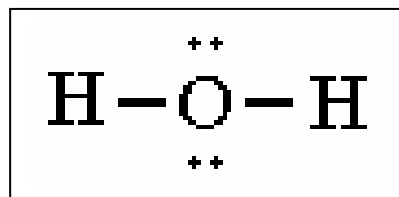
Non-Polar Compound



d) H_2O

P.C. Bonds

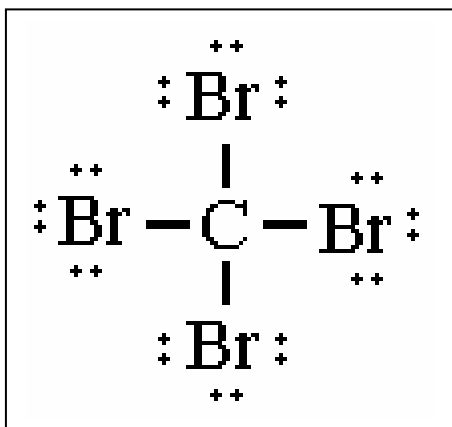
Polar Compound



e) CBr_4

P.C. Bonds

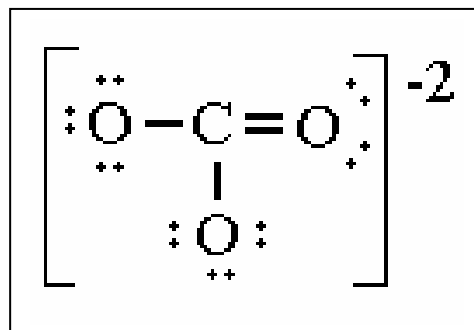
Non-Polar Compound



f) CO_3^{-2}

P.C. Bonds

Non-Polar Compound



g) N_2O

P.C. Bonds

Non-Polar Compound

